

DETERMINATION OF OPTIMAL CONDITIONS FOR EXTRACTING TEREPHTHALIC ACID (H₂BDC) DERIVED FROM PET WASTE USING RESPONSE SURFACE METHODOLOGY (RSM)

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Abstract

This study aimed to develop a practically viable method for extracting terephthalic acid (H₂BDC) from polyethylene terephthalate (PET) waste. Response surface methodology (RSM) was employed to examine the influence and interaction of extractant (NaOH) concentration, hydrolysis time, and size of PET waste on the optimal H₂BDC yield. Experimental results demonstrated that NaOH concentration significantly impacted both H₂BDC yield, following the effects of hydrolysis time and size of PET waste. Under optimal conditions involving a temperature of 160°C and a 12 h reaction time with 15% NaOH, the yield model predicted an 81.58% yield, which closely matched the experimental results of 84.56%. The findings highlight the high PET-derived H₂BDC quality, meeting commercial product criteria with evidence of high purity. This method presents a promising solution for extracting H₂BDC from PET waste, with potential implications for the recycling industry and a positive environmental impact.

Keywords: Terephthalic acid, polyethylene terephthalic (PET), response surface methodology.

JEL Classifications: Q52, Q53, Q55.

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1. INTRODUCTION

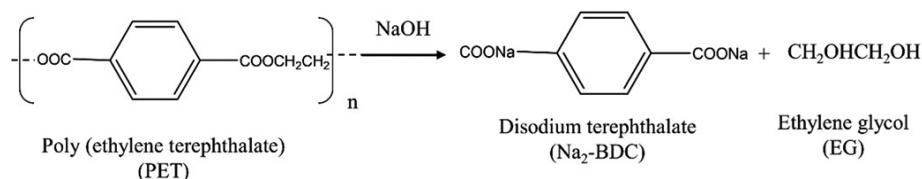
PET bottles have been popularly used for packaging, household goods, clothing, and construction, and have dominated the consumer market over the years. Especially, in developing countries and developed countries, such as Nigeria, Namibia, India, and Vietnam, resulting in the generation of a large amount of PET waste [1, 2]. However, the recycling percentage has remained at a low ratio of only 16%, while the majority are released into landfills and the ocean, leading to a threat to the ecosystem [3, 4]. Significant increase in global plastic manufacture from 1.7 million tonnes in the 1950s to 367 million tonnes by 2020, and is predicted to reach around 590 to 1200 million tonnes by 2050 [2, 5]. Currently, production reaches 414 million tonnes, with PET making up 257 million tonnes of the total, and approximately 177 million tonnes of PET waste are discharged into the environment. Asia is recognized as the world's second-largest producer of plastics, contributing 19.7% of the global total. This analysis underscores the necessity of developing efficient recycling strategies, both to mitigate the environmental impact and to enhance the economic value of plastic waste [6, 7].

Polyethylene terephthalate (PET) contains approximately 85 wt% terephthalic acid (H₂BDC), a highly valuable precursor for the Synthesis of carbonaceous materials and a fundamental linker in the construction of diverse Metal-Organic Frameworks (MOFs) [8, 9]. Accordingly, the development of efficient recycling strategies for PET waste offers

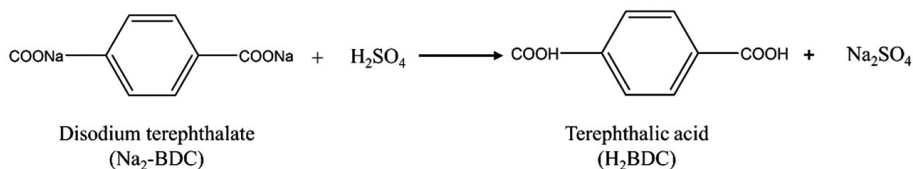
multiple benefits, including extending its material life cycle, alleviating plastic pollution, generating high-value secondary products, and reducing dependence on finite non-renewable feedstocks.

During the past decade, mechanical and chemical techniques have been widely used for recycling PET waste. [10-12]. Mechanical recycling has been extensively applied to reprocess PET waste, yielding secondary products including textile fibers, strapping materials, and regenerated beverage bottles. [5, 13, 14]. Nevertheless, this approach demands elevated processing temperatures (250-280°C) and typically involves labor-intensive sorting and repeated reprocessing, which ultimately diminishes the quality and economic value of the final products [15]. Chemical recycling methods present a promising alternative to mechanical approaches, as they tackle the challenge of plastic recycling by decomposing polymers into their basic chemical constituents [11, 16-18]. The chemical approach fully decomposes PET polymer waste into its monomeric units, which can be repolymerized into high-purity PET, thereby enabling its use in high-performance applications [19]. Exhibiting high efficiency under relatively low conditions of temperature and pressure, this approach offers a cost-effective and energy-efficient pathway for PET degradation [11, 20]. Chemical recycling provides a more sustainable and economically practical strategy for tackling issues related to plastic recycling [16, 21].

The chemical approach includes various techniques such as hydrolysis, glycolysis, aminolysis, and



Scheme 1. PET hydrolysis process using alkaline



Scheme 2. Acidification process using sulfuric acid

methanolysis [22]. Hydrolysis represents a highly effective approach for PET conversion, decomposing PET waste into terephthalate salts and ethylene glycol (EG) [14, 16, 20] (Scheme 1). The obtained terephthalate salt can subsequently be converted into its monomer, hydrogenated H₂BDC, via an acidification process (Scheme 2). Previous studies have demonstrated that PET can be depolymerized into H₂BDC and EG through hydrolysis, with the resulting H₂BDC subsequently employed for the synthesis of Cr-MOF, which exhibits excellent performance in hydrogen storage [23, 24].

Prior research has predominantly centered on extracting H₂BDC from PET waste via hydrolysis and acidification techniques, showcasing promising yield and recovery ratio results. However, these studies were constrained by the reliance on hydrolysis catalysts. Moreover, previous investigations need to

pay more attention to the main aspects, such as the determination of the optimum conditions for efficient H₂BDC extraction from PET waste. In 2020, Jung et al. identified optimal conditions for hydrolyzing powdered PET utilizing an ultrasonicator method, yielding a substantial H₂BDC output [25]. However, the study did not examine how the particle size of PET waste influences the efficiency of H₂BDC recovery during the hydrolysis process.

In summary, this study presents a cost-effective approach utilizing Response Surface Methodology (RSM), which delivers notable economic and environmental advantages. By optimizing hydrolysis parameters, RSM reduces chemical and solvent consumption, lowers energy demand, and decreases wastewater treatment costs. Accordingly, the objectives of this work are twofold: (1) to explore the interactions among independent variables-including NaOH concentration, PET particle size, and reaction time-using RSM to optimize the hydrolysis process and maximize both H₂BDC yield and recovery efficiency; and (2) to evaluate the purity of PET-derived H₂BDC under the optimized conditions in comparison with that of commercially available H₂BDC.

2. MATERIALS AND METHODS

2.1. Materials

PET waste was collected from household recycling bins in Ho Chi Minh City, natural resources and environment, and subsequently prepared for experimentation. The collected PET was first dried in an oven at 90°C for 24h, then thoroughly washed with tap water. All chemicals used in the experiments, including H₂SO₄ and NaOH, were supplied by Bach Khoa Company. All chemicals were used without the need for further purification. The characteristics of PET waste were thoroughly examined through proximate and ultimate analyses by Taiwan Environmental Protection Administration standards (NIEA R213 and R205). The study included the determination of moisture,

Table 1. Basic characteristics of PET plastic waste

Characteristics	Value
Proximate analysis (wt.%, as received)	
Moisture	0.36 ± 0.05
Ash content	0.02
Volatile matter	94.09 ± 0.51
Fixed carbon	5.53 ± 0.51
Ultimate analysis (wt.%, wet basis)	
H	2.94 ± 0.58
C	54.51 ± 2.91
N	ND
S	ND
Cl	ND
O	42.17
Energy content (Kcal kg ⁻¹)	
Higher heating value	5335
Lower heating value	5316

ND: Not detected

ash content, volatile matter, and fixed carbon, with all measurements performed in triplicate to ensure accuracy and reliability. Ultimate analysis of the PET samples was conducted using an elemental analyzer (Vario MICRO, Elementar). As presented in Table 1, the PET waste exhibited a high carbon content, with a dominant ratio of approximately $54.51 \pm 2.91\%$. The calorific value of the samples was determined using a bomb calorimeter (Parr 1341) in accordance with the NIEA E214.01C protocol. Experiments were conducted at the lab of Ho Chi Minh City University of Natural Resources and Environment.

2.2. The extraction process of H_2BDC derived from PET waste

2.2.1. Alkaline hydrolysis process

The extraction was carried out through a two-step process involving hydrolysis followed by acidification, as depicted in Fig. 1. During the hydrolysis stage, one gram of PET flakes was treated with a solution containing varying concentrations of NaOH (5, 10, and 15%) and distilled water. The mass ratios were fixed at 3:1 for NaOH to PET and 10:1 for distilled water to PET. The hydrolysis reaction was conducted in a 100 mL Teflon-lined stainless-steel autoclave reactor. The reactor was operated at a fixed temperature of 160°C for reaction times of 8, 10, and 12 h. After completion, it was cooled to room temperature. Subsequently, a 20% NaOH solution was introduced to adjust the pH to 14, thereby ensuring the complete dissolution of residual H_2BDC [26]. Subsequently, the unreacted solid phase of PET was separated from the mixture using a filtration.

2.2.2. Acidification process and calculation of H_2BDC production yield

The precipitation efficiency was evaluated through acidification reactions carried out with 4.5 M H_2SO_4 ,

as shown in Fig. 1. Sulfuric acid was incrementally introduced into the filtrate until the pH was lowered to 1.0, following the procedure reported by previous authors [27, 28]. During the experiments, the precipitation reactions were carried out under constant stirring at 250 rpm. The obtained H_2BDC was separated by filtration, rinsed with deionized water, and subsequently dried in an oven at 70°C for 12h [15]. The conversion efficiency of PET and the yield of H_2BDC were then determined. After the precipitation and cooling steps, the acid-containing aqueous phase was separated by filtration and subsequently reused for further precipitation reactions.

PET conversion and H_2BDC yield (%) were determined as follows: Eqs. (1) and (2):

$$\text{PET conversion (\%)} = \frac{(W_{\text{PET},i} - W_{\text{PET},f})}{W_{\text{PET},i}} \times 100\% \quad (1)$$

$W_{\text{PET},i}$ is the initial weight of PET, and $W_{\text{PET},f}$ is the final weight of PET that remains on the filter paper after filtration.

$$\text{Yield (\%)} = \frac{W_{\text{precipitate}}}{W_{\text{theoretical}}} \times 100\% \quad (2)$$

$W_{\text{theoretical}}$ is the weight of H_2BDC recovered theoretically (100 % recovery of 3 g of PET is 2.596 g of H_2BDC [25, 29], and $W_{\text{precipitate}}$ is the weight of H_2BDC extracted from designed experiments (Fig 1).

2.3. Experimental design for hydrolysis reaction using RSM

Response Surface Methodology (RSM) is an important statistical approach for optimizing experimental conditions while reducing the number of required trials. It enables the effective optimization, maximization, or minimization of output responses influenced by multiple variables [30, 31]. RSM is employed to simulate the interactive effects of key input variables on the output response. Among its

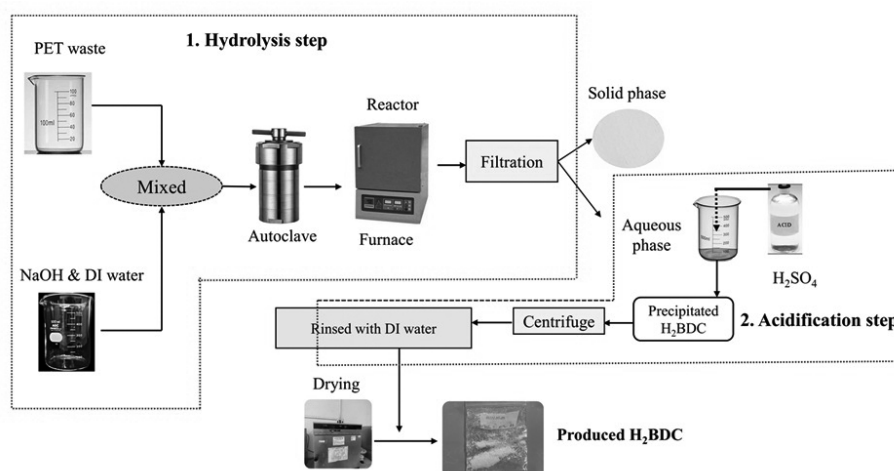


Fig 1. The PET-derived H_2BDC extraction procedure



designs, the Box-Behnken Design (BBD) is particularly valuable for analyzing linear-squares response surfaces and constructing second-order polynomial models. BBD is a spherical, three-level fractional factorial design that incorporates a central point along with midpoints on the edges of a circumscribed circle [32]. In this study, NaOH concentration, PET waste particle size, and reaction time were identified as key factors affecting the extraction efficiency of H₂BDC from PET waste. The Box-Behnken Design (BBD) was applied to establish the optimal hydrolysis conditions using the hydrothermal method in an autoclave. The design consisted of three levels (-1, 0, +1), comprising 15 experimental runs with five replicates at the central point. The independent variables were NaOH concentration, PET particle size, and reaction time, while the response variable was the H₂BDC yield [25]. These values were selected based on previous research findings and initial screening experiments, ensuring thorough coverage of the parameter space [25,33]. Specifically, NaOH concentrations between 5 and 15%, PET waste particle sizes ranging from 2.0 to 12 mm, and hydrolysis durations of 8 to 12 h were examined, as summarized in Table 2. The analytical process was further supported by Minitab software (version 20), which assisted in data interpretation and model construction. This systematic approach streamlined the optimization procedure while improving the robustness of the experimental design. The RSM results were fitted to a second-order polynomial equation, expressed as Eq. (3) [25, 34].

$$Y = \beta_0 + \sum_{i=1}^3 \beta_i X_i + \sum_{i=1}^3 \beta_{ii} X_i^2 + \sum_{i<j}^n \beta_{ij} X_i X_j + \varepsilon \quad (3)$$

where Y is the predicted response, X_i and X_j are independent variables; β_0 is an intercept coefficient, β_i is a linear coefficient, β_{ii} is a linear-squares coefficient, β_{ij} is an interactive coefficient; n is the number of independent variables, and ε is the experimental error. Variance (ANOVA) analysis is conducted to assess the statistical significance of the polynomial model at a 95% confidence level and to reveal the interaction between independent variables and the response variable.

2.4. Characterization

The elemental composition of PET samples and the resulting H₂BDC is assessed using an elemental analyzer (Vario MICRO). The calorific value of PET samples is measured with a bomb calorimeter (Parr 1341, NIEA E214.01C). X-ray diffraction (XRD, Bruker D8 Advance) with CuK α radiation ($\lambda = 1.5406 \text{ \AA}$) is employed to obtain the diffraction patterns of H₂BDC, using a scan rate of $0.03^\circ \text{ s}^{-1}$ across the 2θ range of $5\text{--}80^\circ$. Fourier transform infrared (FTIR) spectra (Frontier MIR/FTIR) are collected in the wavelength range of $650 - 4,000 \text{ cm}^{-1}$ at a resolution of 0.4 cm^{-1} . The morphology of H₂BDC is characterized via field emission-scanning transmission electron microscopy (FE-SEM, Leica Stereoscan 420).

3. RESULTS AND DISCUSSION

3.1. Optimization of the hydrolysis process using the RSM

3.1.1. Development of the regression model

The preliminary results of H₂BDC extraction from PET waste are summarized in Table 2, showing that NaOH concentration, particle size, and reaction time had a significant impact on both PET conversion and H₂BDC yield. The maximum yield of $84.56 \pm 0.79 \%$ was achieved under the conditions of 15% NaOH, 2mm particle size, and a 12 h reaction time at a fixed temperature of 160°C . The results in Table 2 served as a foundation for selecting and designing the experimental conditions of the independent

variables for RSM. Data obtained under the designed experimental conditions were then applied to develop the regression model. The statistical parameters presented in Table 3 revealed that the linear-squares model provided the most accurate prediction of H₂BDC yield, evidenced by a high predicted R-squared value ($R\text{-pred} = 0.89$), a highly significant P-value (0.000), and a low predicted residual error sum of squares (PRESS*) of 148. Furthermore, this model exhibited the highest adjusted determination coefficient ($R\text{-squared adj} = 0.95$) and the lowest PRESS* among all the evaluated models (as shown in Table 3), accounting for 95% of the response variation.

The reliability of the linear-squares model was assessed using statistical parameters and the ANOVA results presented in Table 4. An F-value of 47 indicated significant variation among the group means, confirming the strong significance of the model. The P-value for H₂BDC yield was below 0.0001, while the “Prob > F” values for nearly all terms in Table 4 were less than 0.05, and a small predicted residual error sum of squares (PRESS*) of 148, demonstrating that the linear-squares model and its terms were statistically significant at the 95% confidence level. The adjusted R^2 value of 0.95 and the predicted R^2 value of 0.89 showed no substantial difference, indicating that the linear-squares model effectively describes the experimental data. Additionally, the variance inflation factor (VIF) of 1.00 for all of the model’s factors, obtained from the RSM results using Design of Experiments, confirmed the independence of the variables in the model. From

Table 4, the Lack of Fit F-value of 931 exceeded the significance level of $\alpha = 0.05$, proving that it was not statistically significant. This result suggests there is only a 0.1% probability that the Lack of Fit F-value occurred due to random noise [25]. Overall, the statistical evaluation demonstrates that the linear-squares model is both highly significant and reliable for interpreting the experimental data. Accordingly, the linear-squares regression equation for H_2BDC yield is presented in Eq. (4).

$$\begin{aligned} \text{Yield}(\%) = & -6.1 - 2.45X_1 + 7.54X_2 + 5.37X_3 + 0.098X_1^2 - 0.29X_2^2 \\ & - 0.16X_3^2 \end{aligned} \tag{4}$$

X_1 , X_2 , and X_3 are symbols of particle size (mm), reaction time (h), and NaOH concentration, respectively. Yield (%) is H_2BDC production.

3.1.2. Proposed regression model analysis

The validity of the proposed linear-squares model was further assessed using Fig. 2. The bell-shaped residual histogram in Fig. 2a confirmed the normality assumption of the model. As shown in Table 2, the experimental H_2BDC

yields closely matched the values predicted by the linear-squares regression model, demonstrating strong agreement. These findings indicate that the proposed linear-squares model provides reliable and accurate predictions of the experimental outcomes [35, 36]. Furthermore, Fig. 2b illustrates that the internally studentized residuals versus fitted values remained within the ± 2.0 range, with no abnormal deviations, confirming the suitability of the experimental data for the predicted values derived from Equation (4). The normal probability plot of residuals shown in Fig. 2c was used to assess the assumption of constant variance. The even distribution of data points on both sides of the line supports the validity of the proposed model. In addition, the diagnostic plot closely followed a straight line, indicating no need for data smoothing. Therefore, the proposed linear-squares model is considered reliable for predicting experimental outcomes.

3.1.3. Optimum condition and interactive effects between variables using a Pareto chart and a three-dimensional response surface

In this study, the Box-Behnken Design (BBD) was applied to optimize the independent variables, including particle size, reaction time, and NaOH concentration for maximizing H_2BDC yield. Desirability was employed as the selection criterion for each parameter, ranging from 0 (completely

Table 2. Comparison results between experimental and predicted H_2BDC yield

STT	Particle size (mm)	Reaction time (h)	NaOH (%)	H_2BDC yield (%)		Residual (%)
				Experiment	Prediction	
1	2	8	10	65.42	68.37	-1.58
2	2	10	5	56.99	58.42	0.02
3	2	10	15	82.12	79.30	1.56
4	2	12	10	76.63	75.25	0.21
5	7	8	5	46.13	45.92	1.68
6	7	8	15	68.28	66.80	0.17
7	7	10	10	65.12	65.07	0.14
8	7	10	10	64.99	65.07	0.02
9	7	10	10	64.97	65.07	0.00
10	7	12	5	54.02	52.80	-0.06
11	7	12	15	70.59	73.68	-1.57
12	12	8	10	58.60	57.50	-0.11
13	12	10	5	47.37	47.55	-1.46
14	12	10	15	67.05	68.43	0.08
15	12	12	10	64.65	64.38	1.68
Optimal conditions	2	12	15	84.56 ± 0.79	81.58	-2.98

Table 3. Summary statistics of model fitting for H_2BDC production yield

Source	R-sq(adj)	R-sq(Pred)	F-value	P-value	PRESS*
Linear	0.87	0.79	33	0.000	276
Linear - squares	0.95	0.89	47	0.000	148
Linear - interaction	0.85	0.58	15	0.001	559
Quadratic	0.97	0.82	48	0.000	243

* PRESS: Predicted residual error sum of squares

undesirable) to 1 (highly desirable). As shown in Fig. 3, the optimized conditions of 2 mm particle size, 12 h reaction time, and 15% NaOH resulted in a maximum H₂BDC yield of 82.78%, with a desirability value of 1.00. The model-predicted yield was in close agreement with the experimental result of 84.56%, showing a residual of 2.98 (Table 2). These findings demonstrate that Response Surface Methodology (RSM) is an effective tool for optimizing PET waste hydrolysis.

The 3D response surface plots in Fig. 4 illustrate the effects and interactions of the independent variables on H₂BDC yield. According to the linear-squares response model, the plots indicate that reducing particle size, increasing NaOH concentration, and extending reaction time significantly enhance H₂BDC yield. The study further identified the optimal PET waste-to-water mass ratio as 1:10, consistent with findings from previous research [15, 25]. The Normal Plot of Standardized Effects highlights NaOH concentration as the most influential factor on H₂BDC yield, followed by reaction time and particle size (Fig. 4d). The plot also indicates that the interaction effects among variables were negligible. Under the optimized conditions of 2 mm particle size, 15% NaOH, 12 h reaction time, and a fixed temperature of 160°C, the depolymerization of 2.0 g of PET waste produced 1.47 g of H₂BDC. This process required 6 mL of 5% NaOH solution and 20 mL of deionized water for hydrolysis, followed by 4 mL of 20% NaOH solution for neutralization and 4 mL of H₂SO₄ 4.5 M. Overall, the results confirm that Response Surface Methodology (RSM) is a reliable tool for identifying optimal conditions and assessing variable effects in PET waste hydrolysis.

Table 4. ANOVA of linear-squares model for H₂BDC yield

Source	DF	Adj SS	Adj MS	F-Value	P-Value	VIF	
Model	6	1298	216	47	0.000		
Linear	3	1203	401	86	0.000		
Particle size	1	236	236	51	0.000	1.00	
Time	1	94	94	20	0.002	1.00	Significant
NaOH	1	872	872	188	0.000	1.00	
Square	3	95	32	6.8	0.013		
Particle size	1	22	22	4.8	0.059	1.00	
Time	1	5	5	1.08	0.329	1.00	
NaOH	1	62	62	13.44	0.006	1.00	
Error	8	37	4.6				
Lack-of-Fit	6	37	6.2	931	0.001		Not significant
Pure Error	2	0.01	0.007				
Total	14						

$$\alpha = 0.05$$

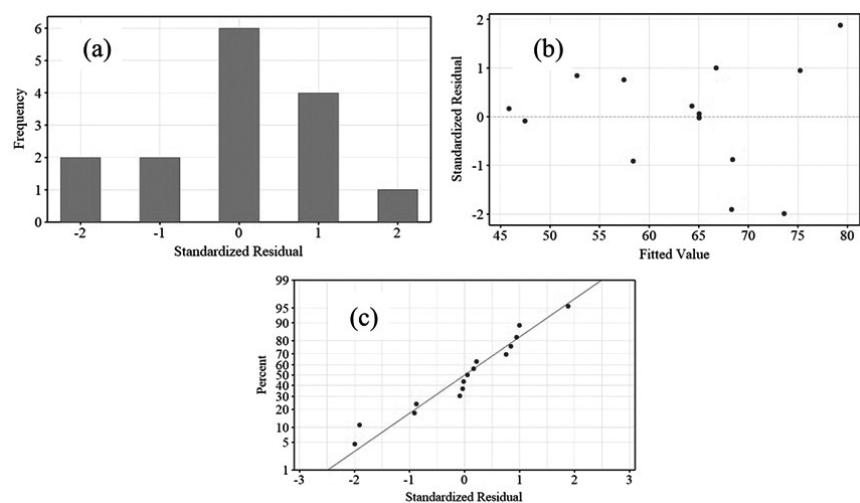
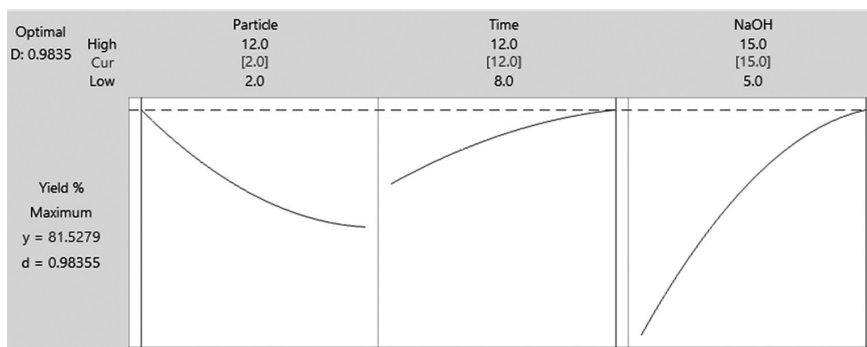


Fig 2: (a) Histogram of residual values and frequency; (b) Standardized Residual and Fitted Value of the linear-squares model; (c) Normal probability and internally studentized residual.



D: Desirability

Fig 3. Optimization of the H₂BDC Yield

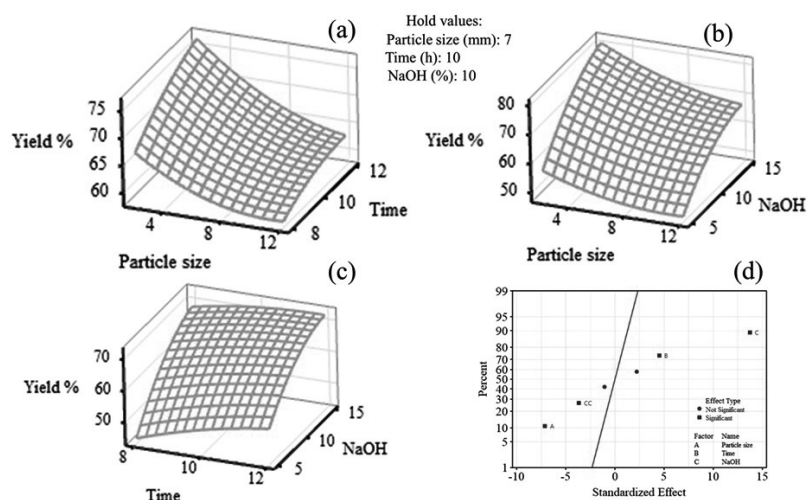


Fig 4. 3D response surface plots for H_2BDC yield (a, b, and c) illustrating the interactive effects: (a) Particle size and time; (b) Particle size and NaOH; (c) Time and NaOH; (d) Normal Plot of the Standardized Effects.

3.2. Evaluation of PET-derived H_2BDC quality at optimum conditions

The PXRD patterns shown in Fig. 5a display several sharp Bragg reflections at 17.2° , 25.3° , 27.9° , 29.8° , and 39.8° , confirming the high crystallinity of PET-derived H_2BDC . The characteristic peaks of PET-derived H_2BDC align closely with those of commercial H_2BDC and agree with findings reported in earlier studies [23, 25, 37]. FTIR analysis was used to characterize the chemical structures of PET waste, PET-derived H_2BDC , and commercial H_2BDC , as shown in Fig. 5b. The absorption bands at $1,578$, $1,511$, and $1,422\text{ cm}^{-1}$ correspond to the vibrational modes of the benzene ring present in

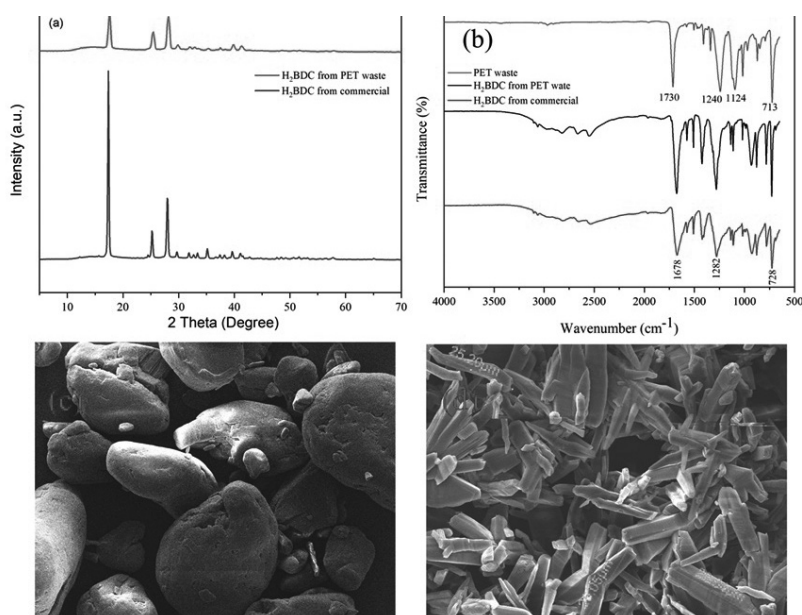


Fig. 5 (a) PXRD pattern of PET-derived H_2BDC vs. commercial H_2BDC ; (b) Comparison of FTIR spectra of PET-derived H_2BDC vs. commercial H_2BDC and original PET. (c-d) FE-SEM image of H_2BDC from commercial and PET waste.

both H_2BDC and PET [25]. Peaks around $1,282$, $2,548$, and $3,062\text{ cm}^{-1}$ correspond to the stretching vibrations of carboxylic groups [16]. The absorption band at $1,678\text{ cm}^{-1}$ corresponds to the stretching vibration of carbonyl groups, whereas the peak at 728 cm^{-1} is associated with the out-of-plane bending vibrations of aromatic rings [25, 38]. In the original PET spectra, the peaks at $1,730$ and $1,252\text{ cm}^{-1}$ are assigned to the $C=O$ and $C-O$ stretching vibrations of ester bonds, respectively. A noteworthy observation in Fig. 5(a–b) is the overlap of characteristic peaks between PET-derived H_2BDC and commercial H_2BDC , validating the successful recovery of H_2BDC from PET bottles. Moreover, the FE-SEM images in Fig. 5(c–d) reveal comparable morphologies, with the PET-derived H_2BDC exhibiting smaller particle sizes than its commercial counterpart. Elemental analysis revealed that the carbon contents of commercial H_2BDC and PET-derived H_2BDC were $58.68 \pm 1.5\%$ and $56.64 \pm 2.2\%$, respectively. Similarly, the hydrogen contents were nearly identical, measured at $3.6 \pm 0.2\%$ for commercial H_2BDC and $3.5 \pm 0.1\%$ for PET-derived H_2BDC . These results confirm that PET-derived H_2BDC exhibits structural consistency, quality, and purity comparable to commercial H_2BDC (99% purity), underscoring its potential for commercial-scale applications. Table 5 summarizes the extraction efficiency of PET-derived H_2BDC obtained under the optimized conditions of this study in comparison with values reported in previous studies [14, 15, 25, 27, 28, 37, 39, 40]. As shown in Table 5, recent studies have focused on catalyst-assisted hydrolysis, which increases operational costs. In contrast, this study identified the optimal hydrolysis and acidification conditions without the use of catalysts, thereby achieving greater cost-efficiency.

4. CONCLUSION

This study determined the optimal hydrothermal conditions for PET hydrolysis, specifically a particle size of

Table 5. Comparison results of PET-derived H₂BDC extraction efficiency

PET size	Agent	T (°C)	Time (h)	Catalyst	Acid	Yield (%)	Ref.
Flakes	NaOH (15%)	160	12	No	H ₂ SO ₄	84.56	This study
Flakes	EG	190	3	No	H ₂ SO ₄	87	[27]
Flakes	EG	210	8	No	ND	55	[37]
Granules	TPA	220	4	TPA	H ₂ SO ₄	95	[15]
Granules	ZSM-5	230	0.5	ZSM-5	HCl	97	[14]
Granules	NaOH (4%)	110	12	CTAB	H ₂ SO ₄	90	[28]
Powder	NaOH (14.5 %)	83.2	1.5	TBAI	H ₂ SO ₄	100	[25]
Powder	NaOH (10%)	90 °C	1	TBAI	HCl	98	[40]

2 mm for 12h with 15% NaOH, with a fixed temperature of 160°C. The process significantly improves the quality of H₂BDC obtained from PET waste while lowering pollutant levels in wastewater through effective acid reuse during acidification and efficient PET recycling. Contrasting previous studies that employed catalysts without optimizing hydrolysis parameters or addressing acid recovery, this work overcomes such limitations. Overall, the results demonstrate strong potential for scaling up H₂BDC production from PET in an environmentally sustainable and economically feasible manner. Future investigations will explore the application of PET-derived H₂BDC in the synthesis of various MOFs and assess their performance in adsorbing and separating pollutants in water, wastewater, and air treatment.

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